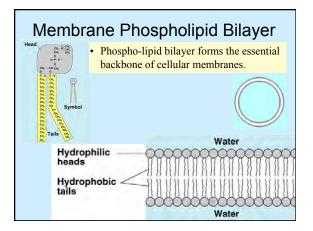


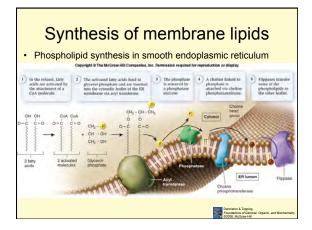
What does a cell need?

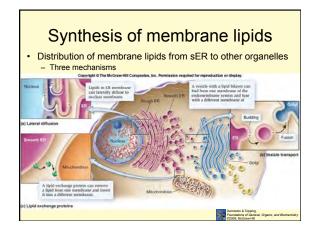
- Selective isolation from environment (plasma membrane)
- Energy (ATP) [to be discussed in future lecture]
- Instructions (DNA)
- Machinery to carry out instructions and regulate processes (proteins)
- Compartmentalization of incompatible or specialized activities (organelles)

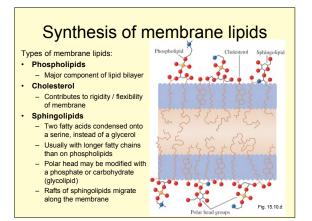
Cell Membrane Function

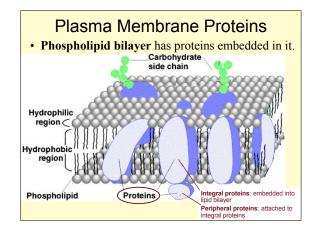
- Boundary between internal and external environments
- Selectively permeable controls what goes in and out of cell or organelle
- Attachment to extracellular surfaces or to other cells (and organelles to cytoskeleton)
- Self or species recognition
- Cell to cell communication
- Lipid metabolism
- · Localization of fixed or sequential processes

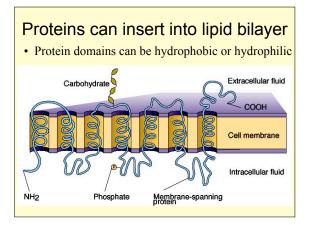


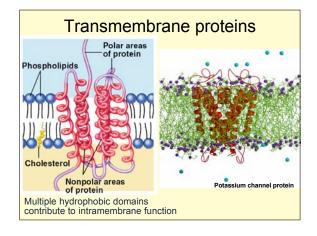


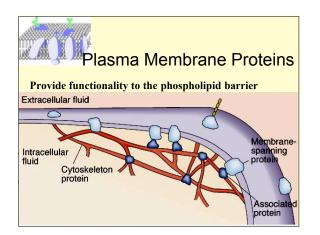


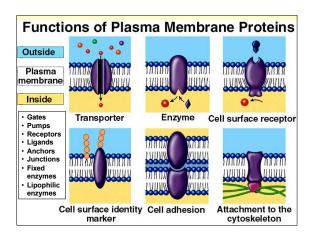




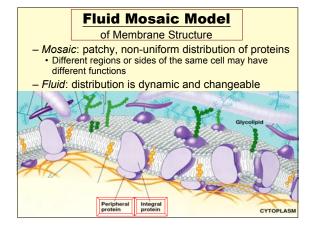


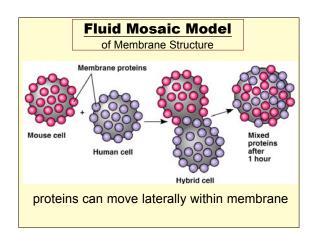


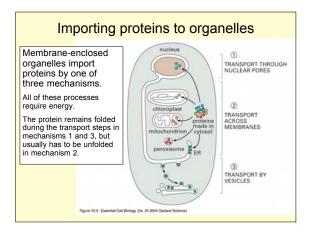


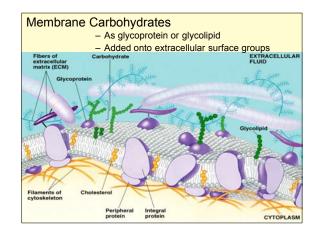


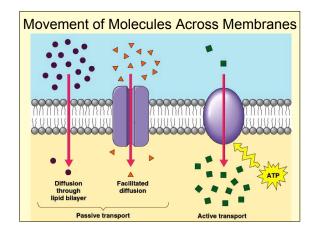
Cell Membranes

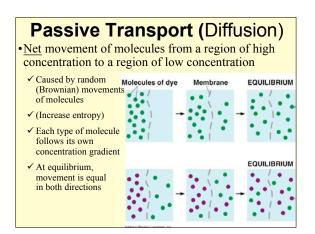


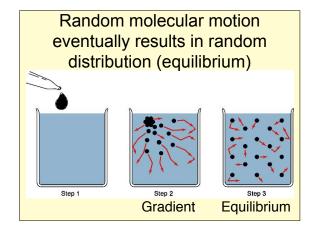








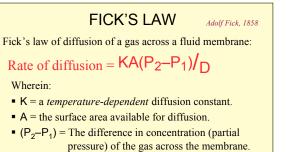




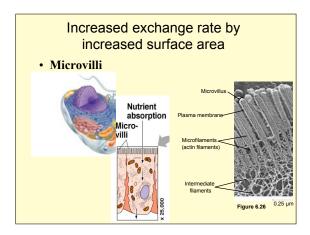
Factors that affect Rate of Diffusion

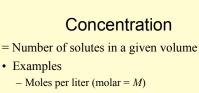


- Concentration gradient
 <u>Difference</u> in concentration between two points
- Temperature (molecular movement)
- Permeability of the membrane / medium
- Available surface area of membrane
- Distance across which diffusion must occur
- Solvent state (gas > liquid > semisolid)

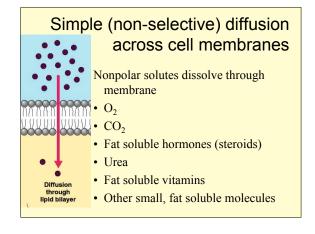


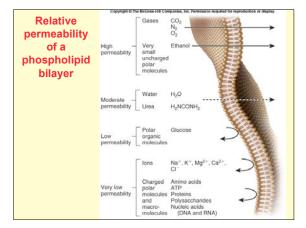
• D = the distance over which diffusion must take place.

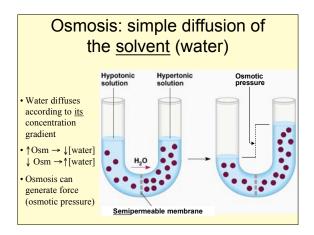




- Grams per 100ml (g%)
- Nanogams per milliliter (ng/ml)
- Parts per thousand (ppt)
- Osmolarity:
- the sum of **all** solutes in a given volume
- in moles per liter (Osm)





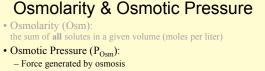


Osmolarity & Osmotic Pressure

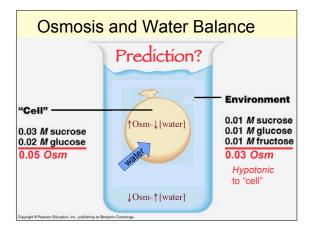
• Osmolarity (Osm):

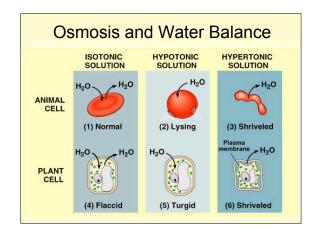
- the sum of **all** solutes in a given volume (moles per liter)
- 1 *M* glucose solution = 1 *Osm*
- 1 *M* glucose/1 *M* fructose/1 *M* ribose solution = 3 *Osm*
- 1 M NaCl solution = 1 M Na⁺/ 1 M Cl⁻ = 2 Osm
- Isosmotic: two solutions with the same Osm
- Hyposmotic: a solution with a lower Osm than another
- Hyperosmotic: a solution with a higher Osm than another

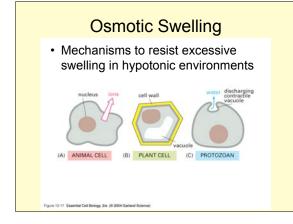
• Remember: $\uparrow \text{Osm} \rightarrow \downarrow [\text{water}]$ $\downarrow \text{Osm} \rightarrow \uparrow [\text{water}]$

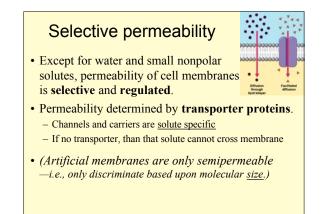


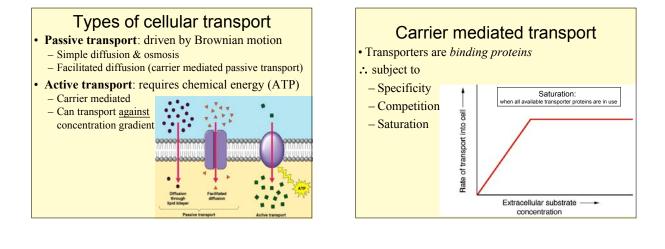
- Measure of the tendency to take on water by osmosis
- Isotonic: two solutions with the same P_{Osm}
- Hypotonic: a solution with a lower P_{Osm} than another I.e., \underline{loses} water to the other solution
- Hypertonic: a solution with a higher P_{Osm} than another I.e., takes water from the other solution
- For an isosmotic solution to be isotonic, the membrane must be equally permeable (or equally impermeable) to all solutes – All isotonic solutions are isosmotic.
 - But not all isosmotic solutions are isotonic.

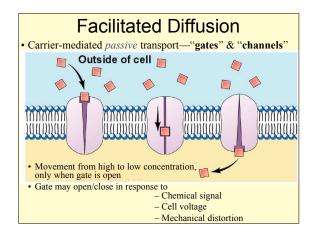


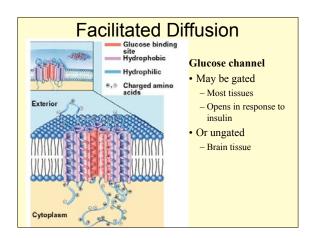


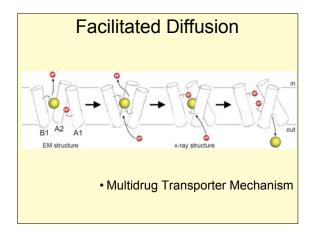


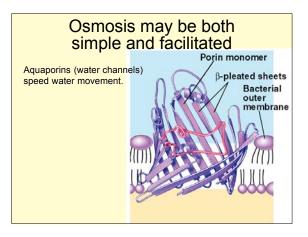


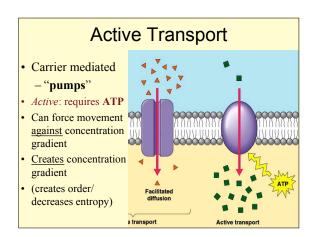


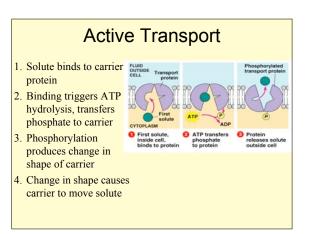


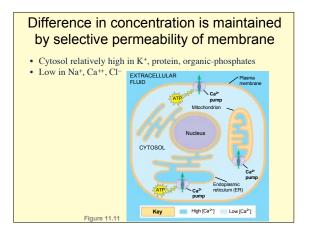


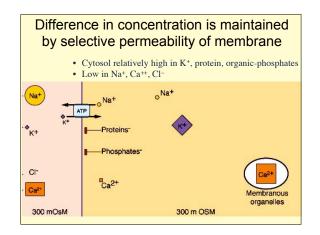


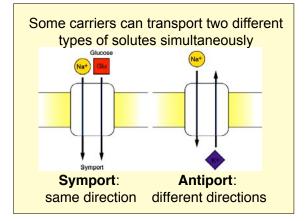


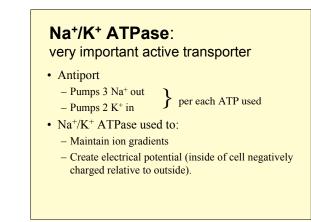


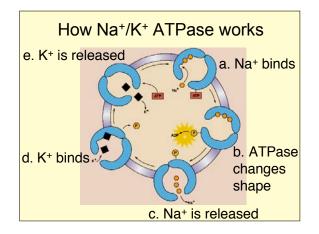


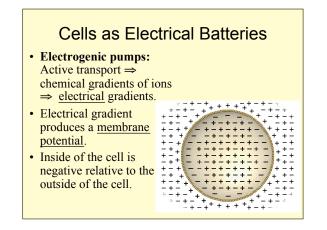


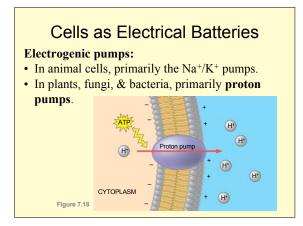


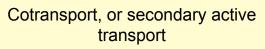












- Carrier protein does not <u>directly</u> use ATP
- But ATP required to create the gradient by other pumps
- Solute "A" transported by diffusion with the created gradient
- Solute "B" moved against gradient by "piggy-backing" with solute "A"
 Outside at cell
- Example: Na⁺ and glucose symport

 Na⁺ diffuses
 - Glucose actively transported



